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Octane isomer dynamics in H-ZSM-5 as a function of Si/Al ratio: a quasielastic neutron scattering study

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Summary

Dynamical behaviour of *n*-octane and 2,5-dimethylhexane in H-ZSM-5 zeolite catalysts of differing Si/Al ratios (15 and 140) was probed using quasielastic neutron scattering, to understand molecular shape and Brønsted acid site density effects on the behaviour of common species in the fluid catalytic cracking (FCC) process, where H-ZSM-5 is an additive catalyst. Between 300 – 400 K, *n*-octane displayed uniaxial rotation around its long axis. However, the population of mobile molecules was larger in H-ZSM-5(140), suggesting the lower acid site concentration allows for more molecules to undergo rotation. The rotational diffusion constants were higher in H-ZSM-5(140) reflecting this increase in freedom. 2,5-dimethylhexane showed qualitative differences in behaviour to *n*-octane with no full molecule rotation, likely due to steric hindrance in the constrictive channels. However, methyl group rotation in the static 2,5-dimethylhexane molecules was observed, with lower mobile fractions in H-ZSM-5(15), suggesting that this rotation is less hindered when fewer Brønsted sites are present. This was further illustrated by the lower activation barrier calculated for methyl rotation in H-ZSM-5(140). We highlight the significant immobilising effect of isomeric branching in this important industrial catalyst, and how compositional changes of the zeolite affect a range of dynamical behaviours of common FCC species upon adsorption.

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Introduction

The behaviour of hydrocarbons in zeolite catalysts is of great interest to a number of applications relevant to the circular economy. One such application is the optimisation/adaptation of processes such as fluid catalytic cracking (FCC) of crude oil fractions to accommodate renewable feedstocks. FCC is one of the major conversion technologies in the petrochemical industry responsible for the majority of the world's gasoline,[1] where ~2300 metric tons of FCC catalyst are produced per day based primarily on zeolite Y[2] but also ZSM-5 to encourage small hydrocarbon products such as propene.[3] The modification of this process to accommodate renewable feedstocks such as biomass[4,5] and even plastic waste[6] requires an in-depth knowledge of how the relevant hydrocarbon species behave in the catalyst micropores.

Of great interest is the effect of molecular shape and branching on the mobility of hydrocarbons, which is fundamental to the shape selectivity characteristics of zeolites.[7,8] The study of this is difficult due to the most important processes taking place under confinement in the porous matrix of the zeolite. However, a significant amount of work both theoretical and experimental has been carried out studying the mobility of hydrocarbons in the MFI zeolite structure[9–16] the framework of H-ZSM-5, and increases in computational power have led to relatively inexpensive computational models of alkane dosed MFI systems reproducing diffusion coefficients that are almost within experimental error.[17–19]

Techniques based on neutron spectroscopy are particularly powerful for studying not only structural and active site characteristics of the catalyst,[20–23] but also the dynamical behaviour of hydrocarbons in microporous catalysts.[24] where quasielastic neutron scattering (QENS) may probe dynamics on the timescale of rotational and translational diffusion.[25] Previous work has employed QENS to study *n*-alkanes in the MFI structure [26] [27] typically yielding diffusion coefficients in the range of $\sim 10^{-10} \text{ m}^2\text{s}^{-1}$. Branched alkanes diffuse over an order of magnitude more slowly[28] and often require higher resolution techniques such as neutron spin echo to gain reliable measurements.[19] Recent work has shown some very revealing insights into localised behaviours, such as the rotation of 1-octene oligomers in the ZSM-5 straight channels[29], and it is clear that such localised motions can bring significant insight into the behaviour of important species as a function of molecular size and shape[30], composition (where the presence of counterions in the framework can significantly affect molecular mobility[26]) and even catalyst use. [31]

To this end, we have employed quasielastic neutron scattering to probe the mobility of common FCC C₈ isomers *n*-octane and 2,5-dimethylhexane in H-ZSM-5, while varying the Si/Al ratio of the catalyst and thus the density of strongly adsorbing Brønsted acid sites. We show that the branching of 2,5-dimethylhexane in

the 5.5 Å ZSM-5 micropores causes significant qualitative differences in motion compared to the straight chain analogue, and that the difference in Brønsted acid site concentration affects the quantitative aspects such as the proportion of mobile populations, and rates of motion observed for both species.

Experimental Section

Materials

The ZSM-5 samples used were commercial zeolite catalysts obtained from Zeolyst International with Si/Al ratio = 15 (CBV3024E) and 140 (CBV28014) referred to herein as H-ZSM-5(15) and H-ZSM-5(140) respectively. Their bulk crystallinity has been verified by powder X-ray diffraction in recent studies.[32,33] The samples arrived in their ammonium form and have been calcined to achieve the catalytic proton form. This was carried out in air, heating at 1 °C min⁻¹ up to 150 °C and 5 °C min⁻¹ to 500°C where they were held for 4 hours. The samples were then dried at 150°C for 4 hours under flowing helium. After cooling, the alkanes were loaded by passing He carrier gas through a Dreschel bottle containing the alkane. The samples were loaded to ~1 molecule of alkane per unit cell, determined gravimetrically.

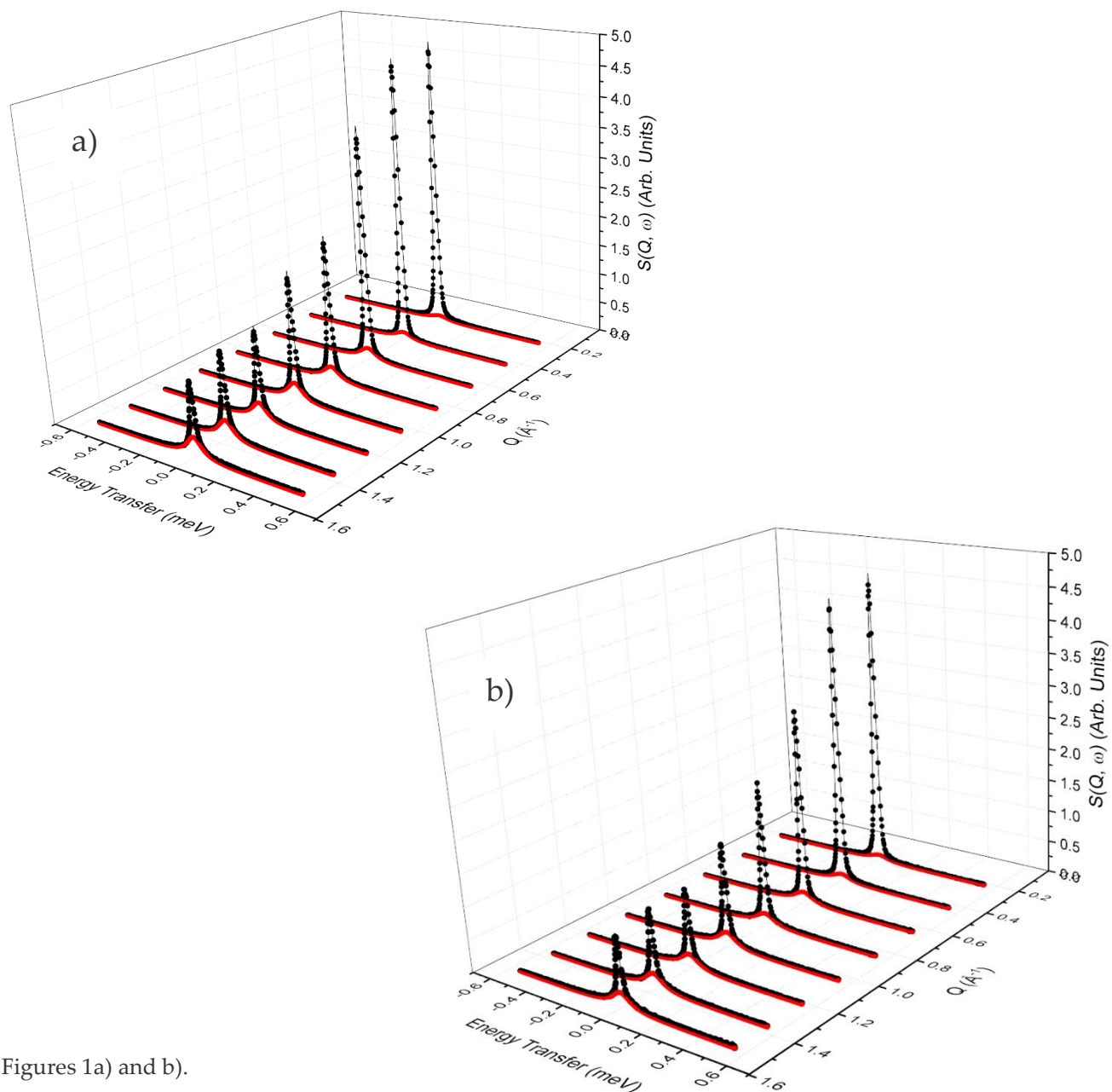
Quasielastic neutron scattering experiments

QENS experiments were carried out using the time-of-flight backscattering neutron spectrometer OSIRIS[34] at the ISIS Pulsed Neutron and Muon Source (UK). The cells were placed in a top loading closed cycle refrigerator. Each sample was then cooled to a base temperature of 10 K and a resolution measurement was taken. The samples were then heated to 300, 360 and 400 K where the QENS spectra were measured. Pyrolytic graphite 002 analyser crystals were used giving an energy resolution of 24.5 µeV with energy transfers measured in a window of ±0.55 meV; the detector coverage enabled measurements over a Q range of 0.2–1.75 Å⁻¹. A measurement was taken of the empty ZSM-5 samples and the signal was then subtracted from the signal of the sorbate loaded ZSM-5, so that the signal from the sorbate only could be extracted. In this way any scattering from the aluminium container, which is very low in comparison with the zeolite is also subtracted. No further corrections were necessary. All QENS spectra were fitted using the neutron scattering analysis software DAVE[35] and MANTID.[36]

Results and Discussion

n-octane dynamics

QENS spectra as a function of the momentum transfer vector Q at 300 K are shown in Figures 1a) and 1b) for *n*-octane in H-ZSM-5(15) and [140] respectively. The spectra were fitted to a delta function convoluted with the resolution measurement taken at 10 K, a single Lorentzian function (which was enough to describe the data satisfactorily) and a flat background function.



Figures 1a) and b).

The figures contain the data points, the total fit (black), and the quasielastic component of the spectra (red) given by the Lorentzian function.

The presence of a large elastic component in all the spectra suggests that either a large fraction of immobile molecules is present or localised motion such as rotation is observed or both. One may characterise the possible localised motions present using the elastic incoherent structure factor (EISF)[37,38] which is given in equation 1, and is the proportion of the total scattered intensity which is elastic.

$$A_0(Q) = \frac{I_{elastic}(Q)}{I_{elastic}(Q) + I_{QENS}(Q)} \quad (1)$$

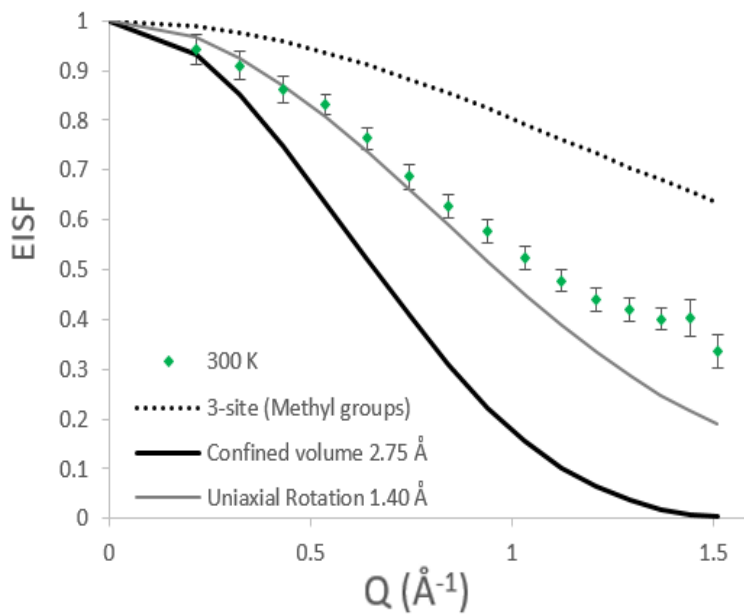


Figure.2.

The experimental EISF for *n*-octane is shown in figure 2, along with models for localised octane motion as outlined in our previous work.[29] These include the Volino and Dianoux model of diffusion confined to sphere, with a radius (r_{conf}) matching that of a ZSM-5 channel, where the EISF is given by equation 2.

$$A_0(Q) = \left[\frac{3j_1(Qr_{conf})}{Qr_{conf}} \right]^2 \quad (2)$$

Where j_1 is the spherical Bessel function of the first kind. The model of 3-site jump rotation around a circle was also employed to represent an octane molecule where only the methyl groups are rotating, given by equation 3 (and shown by the red arrows in figure 3).

$$A_0(Q) = \frac{1}{3} [1 + 2j_0(Qr\sqrt{3})] \quad (3)$$

Where j_0 is the 0th order Bessel function and r in this case is 1.04 Å representing the radius of the circle for the protons of a methyl group, shown by the red arrows in figure 3. The model of uniaxial rotation[30] was employed in our recent work to describe the rotation of oligomerised 1-octene around its long axis (shown by the black arrows in figure 3). The radius of rotation of each octane proton around this long axis (derived in reference 29) is calculated as 1.40 Å (though we note there may be slight deviations with this depending on the behaviour of the terminal methyl groups). This model cannot be used for powder samples typical for porous material studies, because no expression exists for the average angle θ between the axis of rotation and the direction of Q . However, one may consider a jump rotation over N sites around a circle similar to the 3-site model discussed earlier. With a sufficiently large N (> 6) the scattering function does not change as N increases, so the approximation of jump rotation over N sites may then be used as an approximate model for the EISF of continuous rotation, as given in eq. 4

$$A_0(Q) = \frac{1}{N} \sum_{n=1}^N j_0 \left[2Qr_u \sin \left(\frac{n\pi}{N} \right) \right] \quad (4)$$

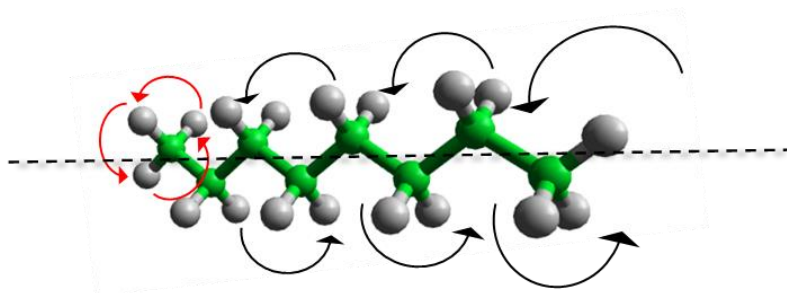


Figure 3.

The model of only methyl rotation falls above the experimental points at all temperatures as shown in figure 2 for the 300 K data (particularly as in this model 2/3 of the protons must be considered static, lifting the model EISF as detailed later), the diffusion confined to a ZSM-5 pore diameter model falls below the experimental points at all but 1 Q value, however, the uniaxial rotation model has a similar shape to the experimental data. We now consider the model EISFs, but with a mobile fraction fitted to take into consideration molecules which are unable to move on the timescale of the instrument, either due to interactions with the catalyst active sites or steric hindrance in the constrictive H-ZSM-5 channels. This is calculated as:

$$A_{0_eff}(Q) = p_x A_0(Q) + (1 - p_x) \quad (6)$$

Where p_x is the fraction of mobile molecules.

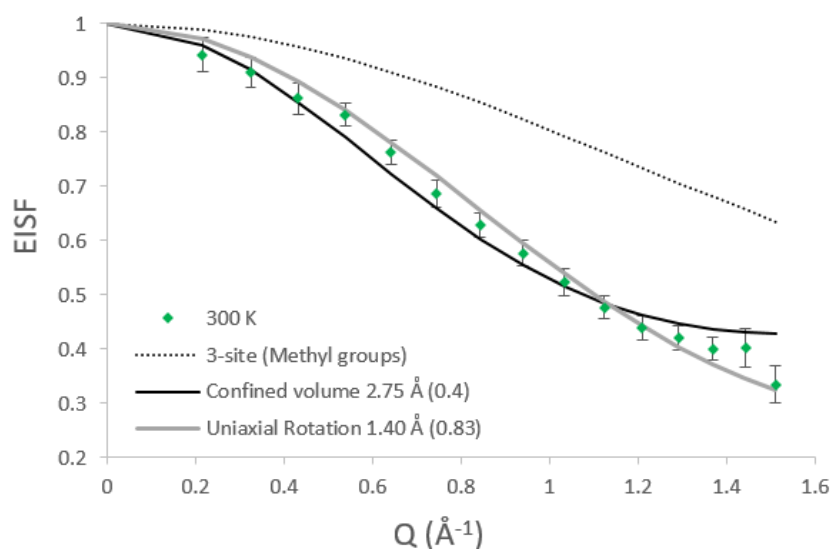


Figure 4.

The optimal fits are shown in figure 4. The best fitting model is that of uniaxial rotation (with 17% of the protons treated as immobile) suggesting the same mode of motion previously reported for oligomeric 1-octene is present for *n*-octane, despite the increased degrees of freedom allowed for the single octane molecule. The error in this fitting is calculated by finding the best fit to the highest and lower points on the error bars, giving an error of $\pm 2.8\%$. This model of motion fits to the higher temperatures in H-ZSM-5(15) as shown in in figure 4 (a) with lower immobile populations (12 and 7 % for 360 and 400 K respectively), and also for all the experimental EISFs in H-ZSM-5(140) with immobile fractions of 11, 6 and 2% at 300, 360 and 400 K

respectively (the error in these values are shown in table 1). We note that the immobile populations are consistently lower in the H-ZSM-5 catalyst with fewer Brønsted sites, suggesting the absence of this 5interaction allows for more molecules to undergo full-molecule uniaxial rotation in the channels. However, we note that the errors bars for each fit to the EISF plots may signify that this difference in immobile populations may be smaller and such differences should be treated with caution.

The magnitude of the quasielastic broadening in the measured neutron spectra is related to the rates of rotation. With rotational motions such as these, the quasielastic broadening will be independent of Q , as shown in figures 5 (a) and (b), where the half-width at half-maxima (HWHM, Γ) of the Lorentzian components fit to the QENS spectra as a function of Q are plotted for all temperatures of n-octane in H-ZSM-5(15) and H-ZSM-5(140), respectively. The broadenings show no Q -dependence, supporting the characterisation of rotational motions, and increase with temperature suggesting an increase in rotational rate as would be expected. We note the widths are larger in ZSM-5[140] suggesting that in this sample not only are more molecules free to rotate, but those that are free also rotate faster in the sample with fewer acid sites.

A model which approximates uniaxial rotation to a jump model of 6 sites around a circle (similarly to how we have modelled the EISF) which was used and detailed by Jobic et al.[30] to describe ethane, propane and other alkanes up to hexane[39] relates the Lorentzian HWHM to the jump rotation as

$$\Gamma = \frac{2}{\tau} \sin^2 \frac{\pi m}{6} \quad (6)$$

Where τ is the mean time between successive rotational jumps (m is the order of the quasielastic structure factor) and the HWHM value at low Q may be used to calculate the rotational diffusion constant D_r as $\frac{1}{2\tau}$ as detailed in reference 30.

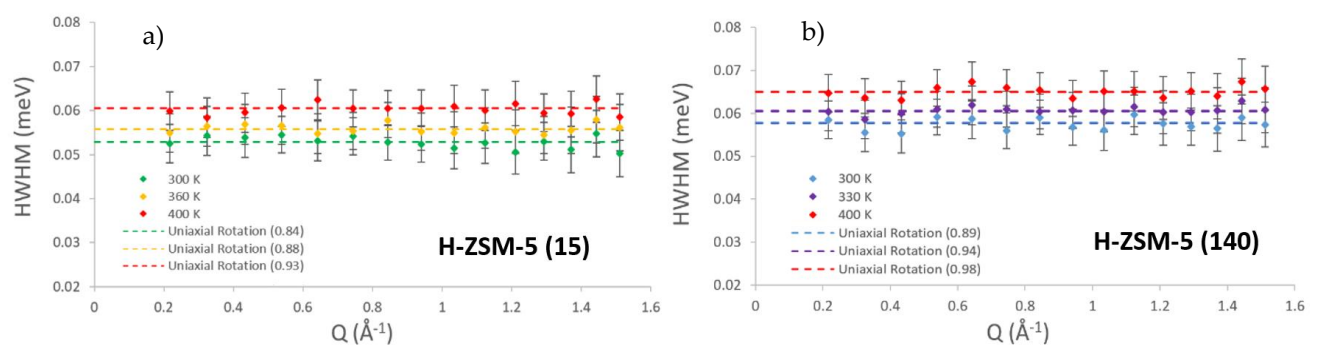


Figure 5 a) and b)

The calculated rotational diffusion coefficients (and mobile fractions) are shown in table 1, ranging from $7.4\text{--}8.5 \times 10^{10} \text{ s}^{-1}$ in H-ZSM-5(15) and $8.1\text{--}9.1 \times 10^{10} \text{ s}^{-1}$ in H-ZSM-5(140), which along with the lower activation energy for rotation, illustrates the increased rotational freedom in ZSM-5 [140]. The D_r values are significantly higher (by about 30%) than those obtained for oligomerised 1-octene in H-ZSM-5 with a Si/Al=30.[29] This is not particularly surprising as the oligomerised chains ($\sim\text{C}_{90}$) are far less mobile than the C_8 monomer analogues in this study, further illustrated by the increase in the mobile population of octane molecules from 29% for oligomerised 1-octene to 88% for *n*-octane $\sim 330 \text{ K}$. The size of the increase in D_r at each temperature between H-ZSM-5(15) and [140] should be treated with caution when the experimental errors are considered. We note that the D_r values are a factor of ~ 5 lower than those calculated for the uniaxial rotation of propane in Na-ZSM-5, even though the loadings are double in terms of carbon atoms per unit cell and the Na^+ ion is present, illustrating the hindering effect of chain length on rotational motion.

2,5-dimethylhexane dynamics

We now discuss the dynamics of 2,5-dimethylhexane in the H-ZSM-5 catalysts. QENS spectra as a function of Q for *n*-octane in ZSM-5[15] and [140] are shown in figures 6 a) and b) respectively. As with *n*-octane, the spectra were fit with to the resolution function, a single Lorentzian function and a flat background function (apart from the lowest 2 Q groups in ZSM-5[15] where only a flat background was necessary, since there was no appreciable Lorentzian component). Upon visual comparison of figures 1 and 6, it is clear that the elastic components are more intense at high Q and the quasielastic components are lower in intensity than for *n*-octane in both samples. In addition, no quasielastic broadening is observed for the lowest 2 groups in ZSM-5 [15] suggesting that no motions are observed at this Q range within the time window of the instrument (1-100 ps), so clearly the movement is very restricted.

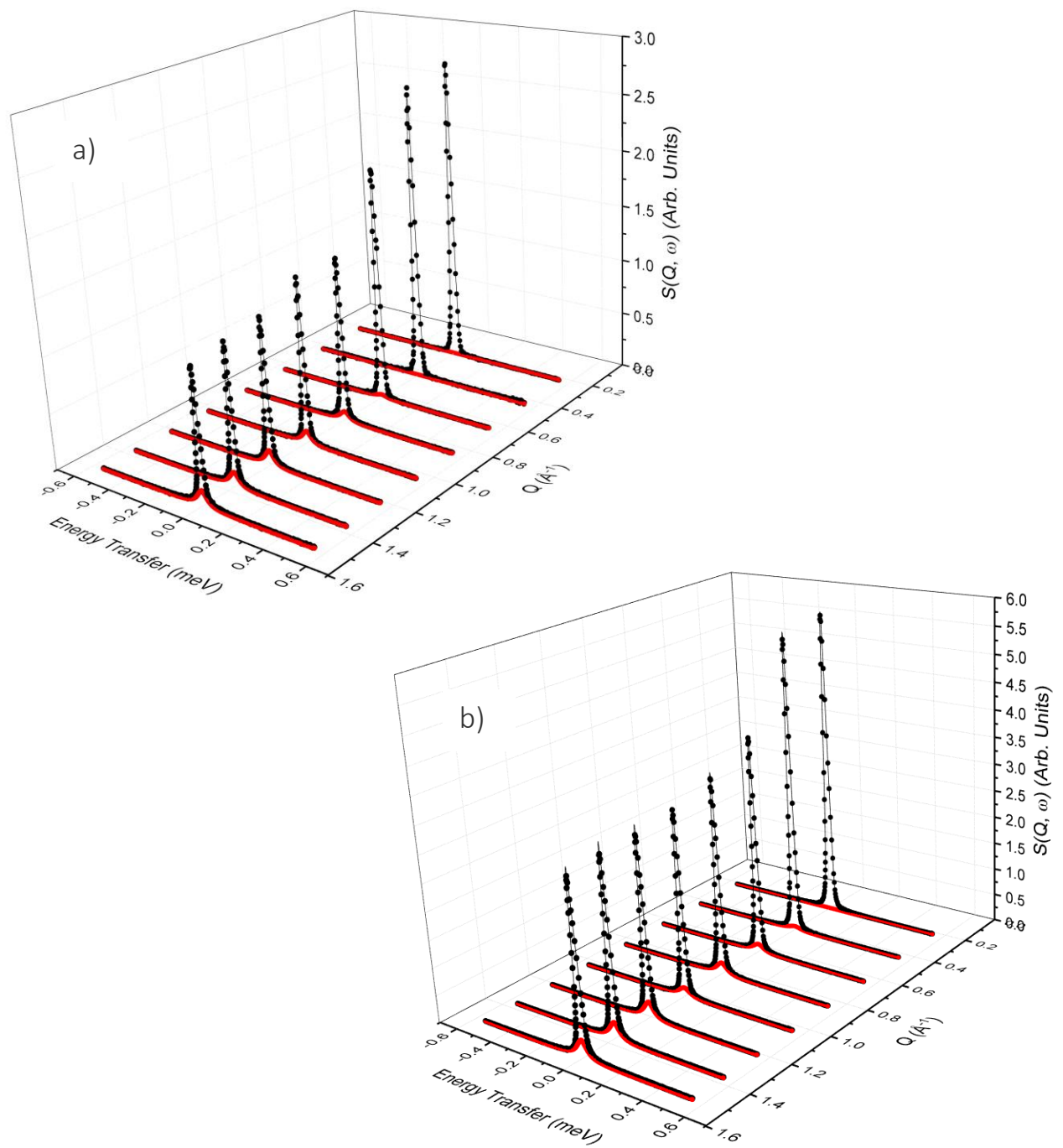


Figure 6 a) and b)

The experimental EISF for 2,5-dimethylhexane in H-ZSM-5(15) at 300 K is shown in figure 7 along with relevant theoretical models, with a mobile fraction considered as for *n*-octane. Along with the models of methyl rotation and diffusion confined to a volume of the ZSM-5 channels we have included a model of uniaxial rotation which takes into account the average radius of all protons from a long axis which passes through the centre of mass of the molecule – such that average radius of rotation (1.75 Å in a 2,5-dimethylhexane molecule geometry optimised using the OPLS[40] forcefield) is feasible in a ZSM-5 channel or intersection, as depicted in figure 8.

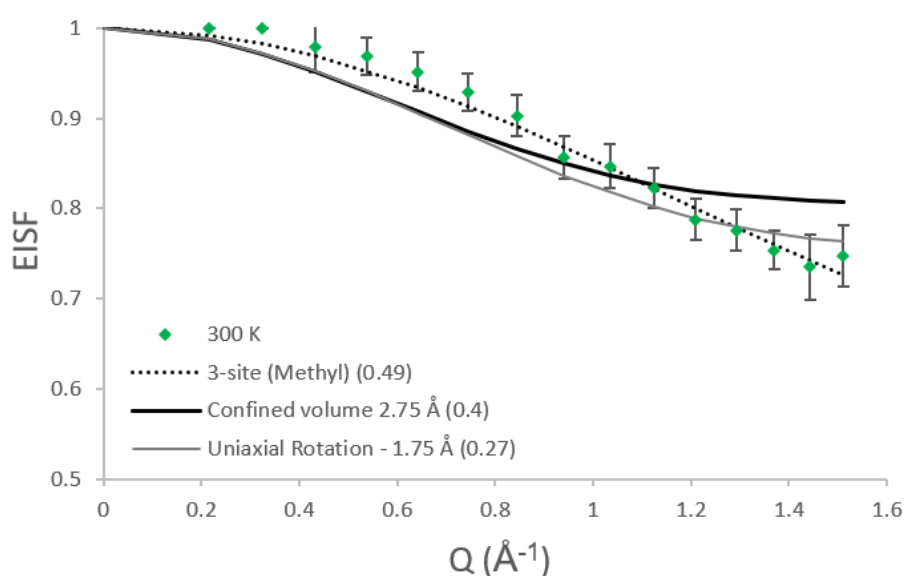


Figure 7.

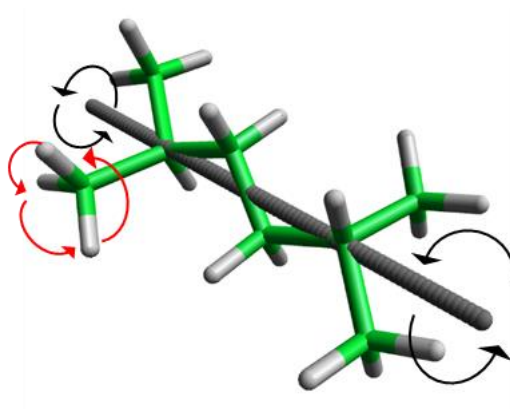
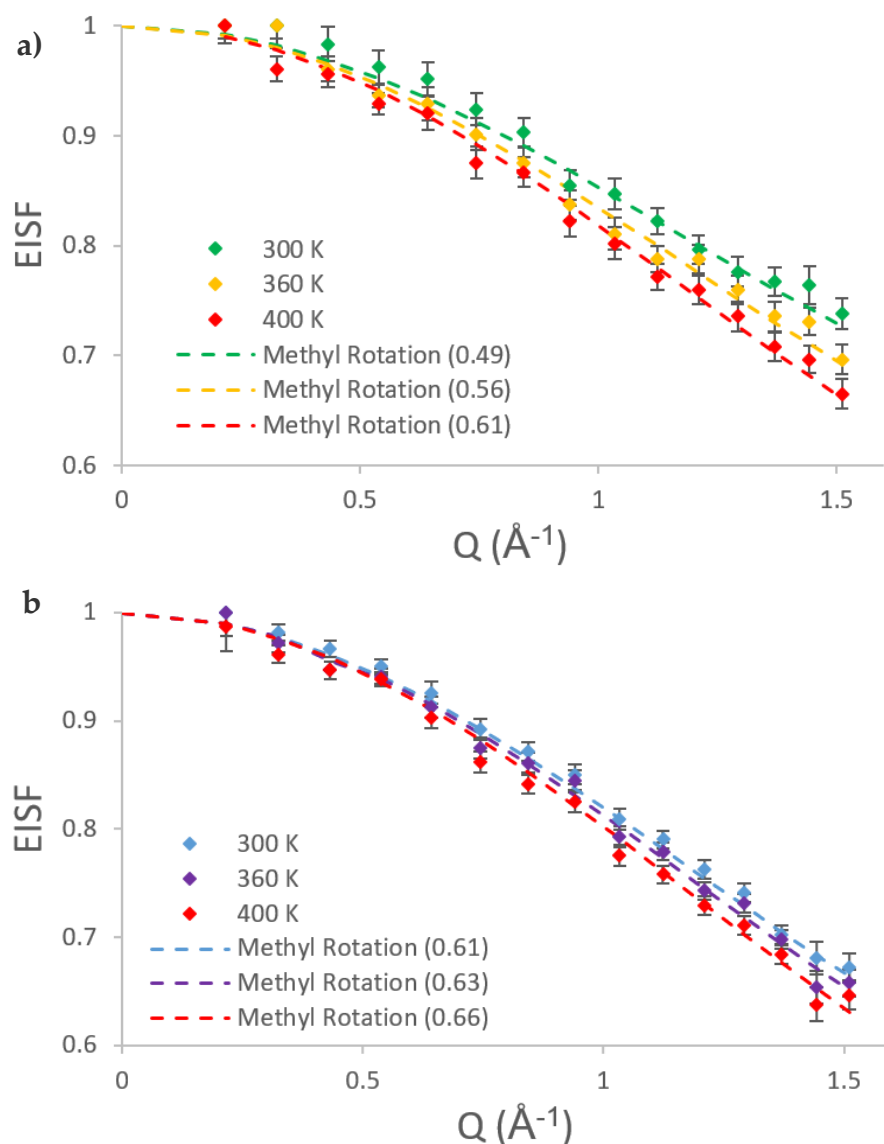


Figure 8.



Figures 9 a) and b).

Perhaps unsurprisingly, given the lack of any translational mobility observed for branched alkanes in previous QENS studies in ZSM-5[28] – the model of confined translation diffusion is not able to fit the data, falling below the experimental points at low Q and above at high Q . The uniaxial rotation model is able to fit the datapoints within error at high Q , but falls below the points at low Q . The 3-site jump rotation model provides the best fit to the data. We note that the highest mobile fraction allowed for this model is $2/3$, as this would encompass all the methyl protons in the molecule rotating, the mobile fraction at this temperature is 0.49, suggesting that in this sample, and at this temperature, movement is so restricted that not even all the

methyl protons are able to rotate in the channels. The model EISF fits the data points for all temperatures in both samples are shown in figures 9 a) and b).

In H-ZSM-5(15) we note that the mobile fraction of methyl protons increases from 0.49 to 0.61 between 300-400 K, whereas in H-ZSM-5(140) the mobile fraction ranges from 0.61 to 0.66. Despite the whole molecule being immobile in terms of both translation and full molecule rotation, we are observing that the lower concentration of Brønsted sites in ZSM-5 allows for more of the methyl groups in the system to rotate, illustrating that interactions with the acid sites can even have a hindering effect even on intramolecular motions. However, we must note that the error bars on the EISF data points are large enough to overlap between temperatures difference by 30 K– so conclusions in terms of mobile fractions must be drawn with caution, though the change in measured EISF is consistent with temperature, and between samples.

As with *n*-octane we may derive information on the rates of rotation from the widths of the Lorentzian component of the QENS spectra. These are plotted in figure 10.

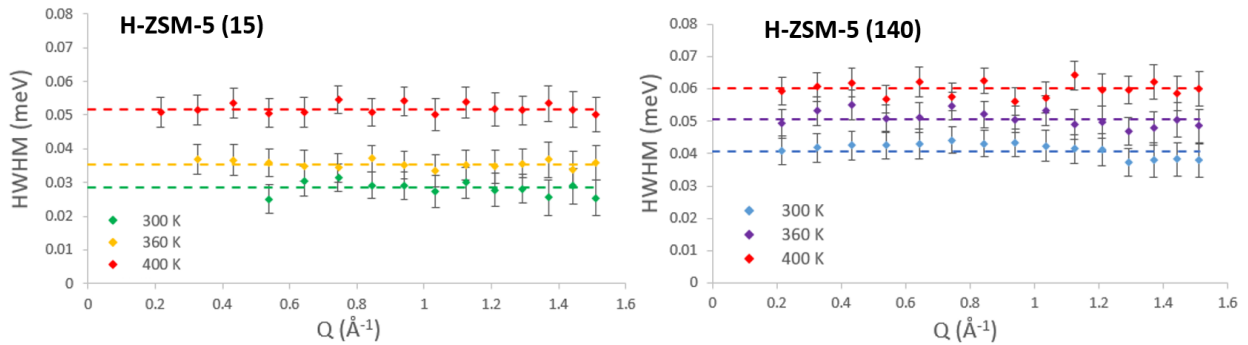


Figure 10

For a model of jump rotation over 3 equally spaced sites around a circle (assuming the probability of each site occupation is equal) the Lorentzian component of the scattering function is given by:

$$L(\omega) = \frac{2\tau/3}{1 + \omega^2(2\tau/3)^2} = \frac{\Gamma}{\Gamma^2 + \omega^2} \quad (6)$$

Where the HWHM Γ is given by $3/2\tau$, and τ is the average time between consecutive jumps. One may calculate the activation energy (the potential barrier hindering rotation) using the Arrhenius law. The residence times and activation energies are shown in table 2, with residence times ranging from 16.5-19 ps in *Phil. Trans. R. Soc. A*.

H-ZSM-5(15) and 15.5-17.5 ps in H-ZSM-5(140). The activation energy to methyl jumps is lower by 0.5 kJ mol⁻¹ in H-ZSM-5(140), suggesting that acidic sites are less significant in the hindering of dimethylhexane methyl rotation than they are in the uniaxial rotation of *n*-octane, and the relatively long residence times are a product of steric hindrance from the constrictive ZSM-5 channels.

On the local motions of hydrocarbons in H-ZSM-5 over the QENS timescale

The introduction of H-ZSM-5 as an additive catalyst to increase the yield of propylene[41] [42] (<5% product yield in typical FCC cracking catalysts[1]) is based on its much narrower pore system than the conventional catalysts based on zeolite Y (~ 5.5 Å intersecting channels in ZSM-5 compared to roughly spherical 13 Å supercages, linked by 7.5 Å windows in zeolite Y). These narrower channels do not allow for bimolecular cracking mechanisms which involve bulkier transition states, and hydrogen transfer reactions. It is this steric constriction and the resulting imposed dynamical restrictions which were the focus of this study, particularly as we expect significant contrast between straight chain and branched isomers, in terms of mobility over the 10⁻¹¹ -10⁻⁹ s timescale probed by QENS instruments. Within this timescale are a range of molecular motions including translational diffusion, local/confined diffusion, full-molecule rotations and intramolecular rotations. Translational diffusion of *n*-octane was probed in ZSM-5 by Jobic et al.[28], obtaining diffusion coefficients in the region of 10⁻¹⁰-10⁻¹¹ m²s⁻¹ using the high resolution backscattering spectrometer IN16 at ILL Grenoble, where the μeV resolution (observing motions in the time window of 0.2-10 ns) allows for the slow diffusive processes in this system to be probed. However, they were not able to determine the diffusivity of a branched alkane (in this case isobutane) as it was too slow even for the higher resolution back scattering spectrometers. Later studies have shown that in the MFI zeolite topology upon which the ZSM-5 is based, the neutron spin-echo technique (which is able to probe motion over the timescale of ~100 ns [25]) can measure diffusivities of such branched species of 1 × 10⁻¹¹ m²s⁻¹ at 444 K [19]– however it should be noted that this is in the fully siliceous MFI system (silicalite-1), without acidic sites/ counterions for the sorbate to bind to which slows diffusion significantly.

It is clear that if one aims to directly compare the motion of straight chain and branched isomers (particularly larger C₈ species) using one instrument, we are limited to more local motions such as rotations, and locally confined translational motions which take place over the faster (ps) timescales where even an order of magnitude difference may be observed in the ~1-100 ps window observed by a backscattering spectrometer with an energy resolution on the order of ~10 μeV. Examples of such rotational studies include the uniaxial rotation of ethane and propane in Na-ZSM-5, where ethane gave D_R values of 1-2 × 10¹¹ s⁻¹ (propane was

similar at 300 K) which is roughly a factor of 2 higher than the uniaxial D_R calculated for octane in this study. One might consider that the difference in rates of motion between ethane/propane and octane in this similar system would be far greater. However, it is important to consider that Na-ZSM-5 (in which ethane and propane were studied) possesses a far larger and more strongly adsorbing counterion than the Brønsted acid sites involved in this study, which would likely hinder rotational motion. A mobile/immobile fraction of molecules is not reported either, and if all the molecules are able to rotate then this system is displaying a far greater degree of rotational freedom than our octane in H-ZSM-5 system as expected. Studies of more bulky molecules in H-ZSM-5 include xylene, [43] where the uniaxial methyl rotation of only one of the two xylene methyl groups was observed, while the other methyl group was said to be interacting strongly with the channel walls, no D_R values were reported for comparison with our dimethylhexane values. However, it is worth noting that the authors gained a lot of insight from the comparison with structural studies and simulations to show how both the orientation and location of *m*- and *p*- xylene isomers lead to this methyl rotation blocking. Cyclohexane rotation in H-ZSM-5 has also been studied[44] and through the fitting of a 3-site jump rotation model it was concluded that cyclohexane was in the chair conformation, the D_{3d} symmetry of which would allow for such rotations, with a residence time of 8.2 ps at room temperature. We note that this is roughly twice as fast as that found for our methyl groups in dimethylhexane which given the size of the cyclohexane molecule/circle around which the protons are jumping may seem strange. However, we must consider that molecules of similar dimensions such as benzene prefer to locate at the channel intersections in ZSM-5, where isotropic and six-fold rotations may be observed.[45,46] Given the much larger void volume of the MFI channel intersections compared to the channels (where we expect our dimethylhexane molecules to be sited if the only motions observable are methyl rotations), the extra freedom of rotation in the larger cyclohexane molecule may not be an unreasonable observation.

It is clear from the referenced studies that probing local motions of hydrocarbons in cracking catalysts can yield some very insightful observations in terms molecular behaviour and interactions, necessary for catalyst optimisation. In the absence of clear vibrational spectroscopic features at the relevant temperatures to illustrate any strong interactions between the catalyst active sites and aliphatic hydrocarbons, such rotational dynamical data may prove a key descriptor for strong interactions with the walls of the catalyst micropores. The value of such local motions in the field of microporous catalysis (certainly in terms of their input into typical kinetic models, where a more standard translational diffusion coefficient is common place) appears yet

to be fully realised, and future QENS/NSE experiments will focus on building a complete picture of dynamical behaviour across scales in addition to these observations.

Conclusion

The dynamical behaviour of *n*-octane and 2,5-dimethylhexane was studied in two H-ZSM-5 samples with Si/Al = 15 and 140 using quasielastic neutron scattering to study the effect of molecular shape, and Brønsted acid site density on the dynamics of common FCC species in acidic zeolite catalysts, where H-ZSM-5 is additive catalyst. Translational diffusion was not observed on the 2-100 ps timescale of the QENS instrument. However for *n*-octane, the EISF was able to be fit to the model uniaxial rotation along its the long axis with rotational diffusion constants in the range of $7.5\text{--}8.5 \times 10^{10} \text{ s}^{-1}$ (with a mobile population of 84-93%) in ZSM-5 [15], and $8.1\text{--}9.1 \times 10^{10} \text{ s}^{-1}$ (with a mobile population of 89-98%) in ZSM-5 [140]. The higher rotational constants and larger mobile populations suggest that the much lower concentration of Brønsted acid sites in ZSM-5 [140] allows for freer rotation of *n*-octane in the ZSM-5 channels. 2,5-dimethylhexane is likely too bulky to fully rotate in the ZSM-5 channels, as the only dynamics which could be observed were the rotations of its methyl groups, which could be fit to a 3-site jump rotation model. The residence times between jumps ranged from 19-16.5 ps (with a mobile population of 49-61%) in H-ZSM-5(15) and 15.5-17.5 ps (with a mobile population of 49-61%) in H-ZSM-5(140). The long residence times reflect the steric hindrance of motion exerted by the ZSM-5 channels, while the (slight) decrease in residence times may be reflective of less hindrance to methyl rotations with fewer Brønsted acid sites present. The study highlights both the effect of molecular shape and branching on the qualitative difference in behaviour of the isomers in H-ZSM-5, while highlighting also the effect the difference in Si/Al ratio (and thus acid site density) on the qualitative rates of motion, and the number of immobile molecules/moieties inside this zeolite catalyst.

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Tables

T K	H-ZSM-5(15)		H-ZSM-5(140)	
	$D_r \text{ s}^{-1}$	Mobile fraction	$D_r \text{ s}^{-1}$	Mobile fraction
300	7.4×10^{10} $\pm 0.25 \times 10^{10}$	0.84 ± 0.023	8.1×10^{10} $\pm 0.27 \times 10^{10}$	0.89 ± 0.021
330	7.8×10^{10} $\pm 0.29 \times 10^{10}$	0.88 ± 0.027	8.5×10^{10} $\pm 0.28 \times 10^{10}$	0.94 ± 0.027
360	8.5×10^{10} $\pm 0.32 \times 10^{10}$	0.93 ± 0.031	9.1×10^{10} $\pm 0.30 \times 10^{10}$	0.98 ± 0.03
$E_a \text{ (kJ mol}^{-1}\text{)}$	4		3.3	

Table 1. Rotational diffusion coefficients (D_{rot}) and mobile fractions of *n*-octane in H-ZSM-5(15) and [140].

T K	H-ZSM-5(15)		H-ZSM-5(140)	
	$\sim \tau \text{ ps}$	Mobile fraction	$\sim \tau \text{ ps}$	Mobile fraction
300	19 ± 0.57	0.49 ± 0.023	17.5 ± 0.55	0.61 ± 0.021
330	18 ± 0.54	0.56 ± 0.027	16.5 ± 0.51	0.63 ± 0.027
360	16.5 ± 0.55	0.61 ± 0.031	15.5 ± 0.53	0.66 ± 0.03
$E_a \text{ (kJ mol}^{-1}\text{)}$	4		3.5	

Table 2. Time between mean jumps (D_{rot}) and mobile fractions and activation energies of methyl jump rotation of 2,5-dimethylhexane in H-ZSM-5(15) and [140].

Figure captions

Fig 1. QENS spectra as a function of Q for n -octane at 300 K in a) H-ZSM-5(15) and b) H-ZSM-5(140). (—) is the total fit to the data points, (—) is the quasielastic Lorentzian component.

Fig.2. Experimental EISF plots of n -octane in ZSM-5 (15) at 300 K against different theoretical EISF models.

Fig.3. Schematic of an n -octane molecule illustrating the methyl rotational motion (red arrows) and the uniaxial rotation of protons around the dashed long axis.

Fig 4. The experimental EISF of n -octane in H-ZSM-5(15) at 300 K, plotted against the models of localised motions after fitting with an immobile fraction. The optimum p_x value is listed in brackets.

Fig. 5. Experimental EISF of n -octane in H-ZSM-5(15) (a) and [140] (b) with the fitted uniaxial rotation model (optimal mobile fractions (p_x) shown in the legend in brackets).

Fig 6. Q -dependence of the HWHM (Γ) of the quasielastic components of QENS spectra of n -octane in each H-ZSM-5 sample from 300-400 K.

Fig 7. QENS spectra as a function of Q for 2,5-dimethylhexane at 300 K in zeolite a) H-ZSM-5(15) and b) H-ZSM-5(140). (—) is the total fit to the data points, (—) is the quasielastic Lorentzian component.

Fig 8. The experimental EISF of 2,5-dimethylhexane in H-ZSM-5(15) at 300 K, plotted against the models of localised motions after fitting with a mobile fraction. The optimum p_x value is listed in brackets.

Fig.9. Schematic of an 2,5-dimethylhexane molecule illustrating the methyl rotational motion (red arrows) and the full molecule uniaxial rotation of protons around the dark grey long axis.

Fig. 10. Experimental EISF of 2,5-dimethylhexane in (a) H-ZSM-5(15) and (b) [140] with the fitted 3-site jump rotation model for methyl rotation (optimal p_x shown in the legend in brackets).

Fig 11. Q -dependence of the HWHM (Γ) of the quasielastic components of QENS spectra of 2,5-dimethylhexane in each H-ZSM-5 sample from 300-400 K.

Authors' Contributions

Experiments carried out or facilitated by AJOM, ND, WJ, SFP and VGS. Analysis carried out by AJOM and VGS. All aspects of the study carried out by AJOM were carried out under the supervision of CRAC. All authors have contributed to the manuscript. All authors read and approved the manuscript.

Competing Interests

The authors declare that they have no competing interest.